



CANDIDATE
NAME

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CENTRE
NUMBER

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CANDIDATE
NUMBER

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0620/32

October/November 2020

1 hour 15 minutes

You must answer on the question paper.

No additional materials are needed.

- Answer **all** questions.
- Use a black or dark blue pen. You may use an HB pencil for any diagrams or graphs.
- Write your name, centre number and candidate number in the boxes at the top of the page.
- Write your answer to each question in the space provided.
- Do **not** use an erasable pen or correction fluid.
- Do **not** write on any bar codes.
- You may use a calculator.
- You should show all your working and use appropriate units.

- The total mark for this paper is 80.
- The number of marks for each question or part question is shown in brackets [].
- The Periodic Table is printed in the question paper.

This document has **20** pages. Blank pages are indicated.

1 (a) The diagram shows part of the Periodic Table.

[illegible]

Answer the following questions using only the symbols of the elements in the diagram. Each symbol may be used once, more than once or not at all.

State the symbol of the element that:

- (i) provides an inert atmosphere in lamps

- (ii) forms an oxide which is used to neutralise acidic industrial waste

..... [1]

- (iii) has an atom which forms a stable ion by the loss of one electron

- (iv) is a metal used as an inert electrode

..... [1]

- (v) forms an ion whose aqueous solution gives a green precipitate on addition of aqueous sodium hydroxide.

..... [1]

(b) Chlorine is an element.

(i) State the meaning of the term *element*.

.....
 [1]

(ii) An isotope of chlorine is shown.

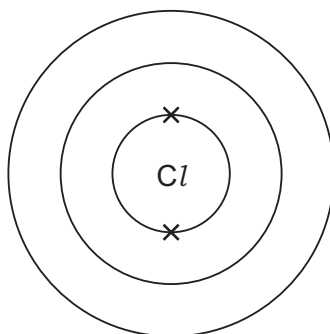


Deduce the number of protons and neutrons in this isotope.

number of protons

number of neutrons [2]

(c) Complete the electronic structure of a chlorine atom.



[1]

[Total: 9]

- 2 The table shows the mass of air pollutants, in nanograms, in 1000 cm^3 samples of air taken over a four month period.

	mass of pollutant in 1000 cm^3 of air/nanograms				
month	oxides of nitrogen	sulfur dioxide	carbon monoxide	ozone	particulates
April	144.3	5.9	2.5	33.9	21.9
May	114.2	2.0	2.1	39.6	21.7
June	110.2	6.1	1.8	31.5	21.3
July	115.4	2.5	2.6	24.2	19.0

(a) Answer these questions using only the information in the table.

- (i) Name the pollutant that shows a continual decrease in concentration between April and July.

..... [1]

- (ii) Name the pollutant present in the lowest concentration in May.

..... [1]

- (iii) Calculate the mass of carbon monoxide in 200 cm^3 of the sample of air taken in April.

..... nanograms [1]

(b) Sulfur dioxide contributes to acid rain.

- (i) State **one** source of the sulfur dioxide in the air.

..... [1]

- (ii) Give **one** adverse effect of acid rain on buildings.

..... [1]

- (iii) State **one** use of sulfur dioxide.

..... [1]

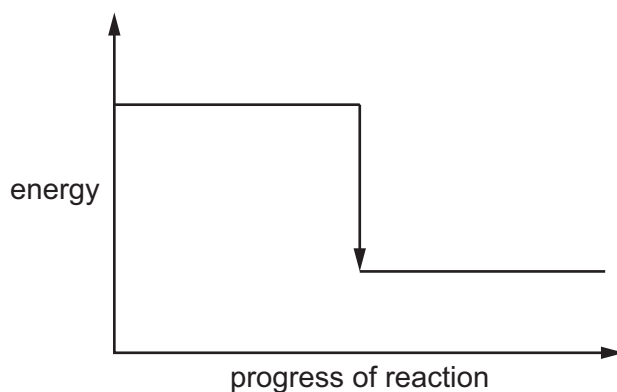
(c) Sulfur dioxide is oxidised to sulfur trioxide, SO_3 .

(i) Complete the chemical equation for this reaction.



(ii) Complete the energy level diagram for the oxidation of sulfur dioxide to sulfur trioxide by writing these words on the diagram:

- reactants
- products.



[1]

(iii) Explain, using information on the energy level diagram, how you know that this reaction is exothermic.

..... [1]

(d) Nitrogen monoxide is a catalyst in the oxidation of sulfur dioxide to sulfur trioxide.

State the meaning of the term *catalyst*.

..... [1]

(e) Sulfur trioxide reacts with water to form dilute sulfuric acid.

Identify which **one** of these pH values represents the pH of dilute sulfuric acid.

Draw a circle around the correct answer.

pH 2

pH 7

pH 9

pH 13

[1]

(f) Particulates are tiny solid particles in the air.

They show Brownian motion.

Identify one statement that best describes Brownian motion.

Tick **one** box.

The particles move from a higher concentration to a lower concentration.

☐

The particles are smaller than oxygen molecules.

☐

Brownian motion is an example of diffusion.

☐

The particles move in a random zig-zag motion.

☐

[1]

[Total: 13]

- 3 Some properties of four substances, **E**, **F**, **G** and **H**, are shown in the table.

substance	strength	ductility (how easy it is to pull into a wire)	electrical conductivity when solid	resistance to corrosion
E	very strong	good	good	very good
F	weak	good	good	poor
G	strong	not ductile	good	poor
H	strong	very good	very good	good

Answer these questions using only the information in the table.

- (a) State which substance, **E**, **F**, **G** or **H**, is best used to make electricity cables.

Explain your answer.

substance

explanation

[3]

- (b) State which substance, **E**, **F**, **G** or **H**, is best used for making cutlery.

Explain your answer.

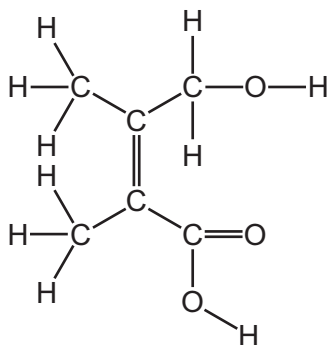
substance

explanation

[3]

[Total: 6]

4 The structure of compound **J** is shown.



(a) (i) On the structure, draw a circle around the carboxylic acid functional group. [1]

(ii) Deduce the formula of compound **J** to show the number of carbon, hydrogen and oxygen atoms.

..... [1]

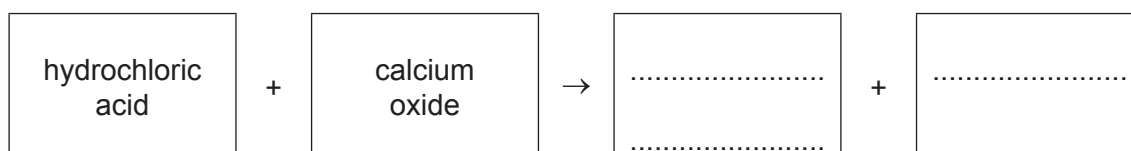
(iii) Complete the table to calculate the relative molecular mass of compound **J**. Use your Periodic Table to help you.

type of atom	number of atoms	relative atomic mass	
carbon		12	
hydrogen	10	1	$10 \times 1 = 10$
oxygen		16	

relative molecular mass = [2]

(b) Acids react with bases such as calcium oxide.

Complete the word equation for the reaction of hydrochloric acid with calcium oxide.



[2]

(c) The chemical equation for the reaction of lime (calcium oxide) with ammonium sulfate is shown.



(i) Name the compound with the formula CaSO_4 .

..... [1]

(ii) Complete these phrases about ammonia, NH_3 , using words from the list.

acid blue gaseous green liquid
pink solid solution white

The state of ammonia at room temperature is

Aqueous ammonia turns damp red litmus paper

[2]

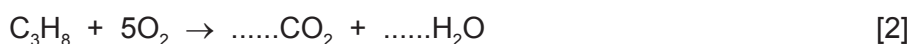
[Total: 9]

5 Ethane is an alkane.

(a) Draw the structure of ethane to show all of the atoms and all of the bonds.

[1]

(b) Complete the chemical equation for the complete combustion of propane.



(c) Methane is an alkane which is produced by the fractional distillation of petroleum.

(i) State one **other** process which puts methane into the atmosphere.

..... [1]

(ii) Give **one** major use of methane.

..... [1]

(d) Alkanes and alkenes are hydrocarbons.

State the meaning of the term *hydrocarbon*.

.....
 [2]

(e) Alkanes and alkenes can be distinguished by a chemical test.

Name the reagent that can be used to distinguish between alkanes and alkenes.

..... [1]

(f) Alkenes are manufactured by cracking alkanes.

(i) Name an element that is also produced by cracking alkanes.

..... [1]

(ii) State **one** condition required for cracking alkanes.

..... [1]

[Total: 10]

6 Electrolysis is used to extract reactive metals from metal compounds.

(a) Describe the electrolysis of molten sodium chloride.

In your answer include:

- a labelled diagram of the apparatus used
- the observations made at the positive and the negative electrode.

observation at positive electrode

.....

observation at negative electrode

.....

[5]

(b) Use the kinetic particle model to describe the arrangement and separation of the particles in solid sodium.

arrangement

separation

[2]

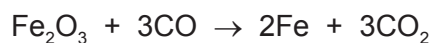
- (c) Sodium is a metal in Group I of the Periodic Table.
Iron is a transition element.

Give **two** ways in which the physical properties of iron differ from the physical properties of sodium.

1

2 [2]

- (d) The chemical equation for the reaction between iron(III) oxide and carbon monoxide is shown.

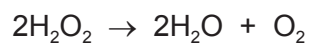


Explain how this equation shows that carbon monoxide has been oxidised.

..... [1]

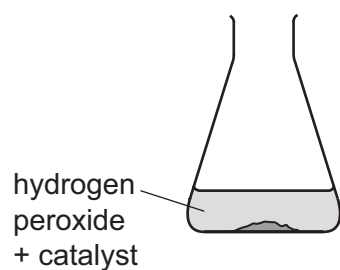
[Total: 10]

- 7 A student investigated the rate of decomposition of hydrogen peroxide, H_2O_2 , in the presence of a catalyst by measuring the volume of oxygen released at 10 second intervals.



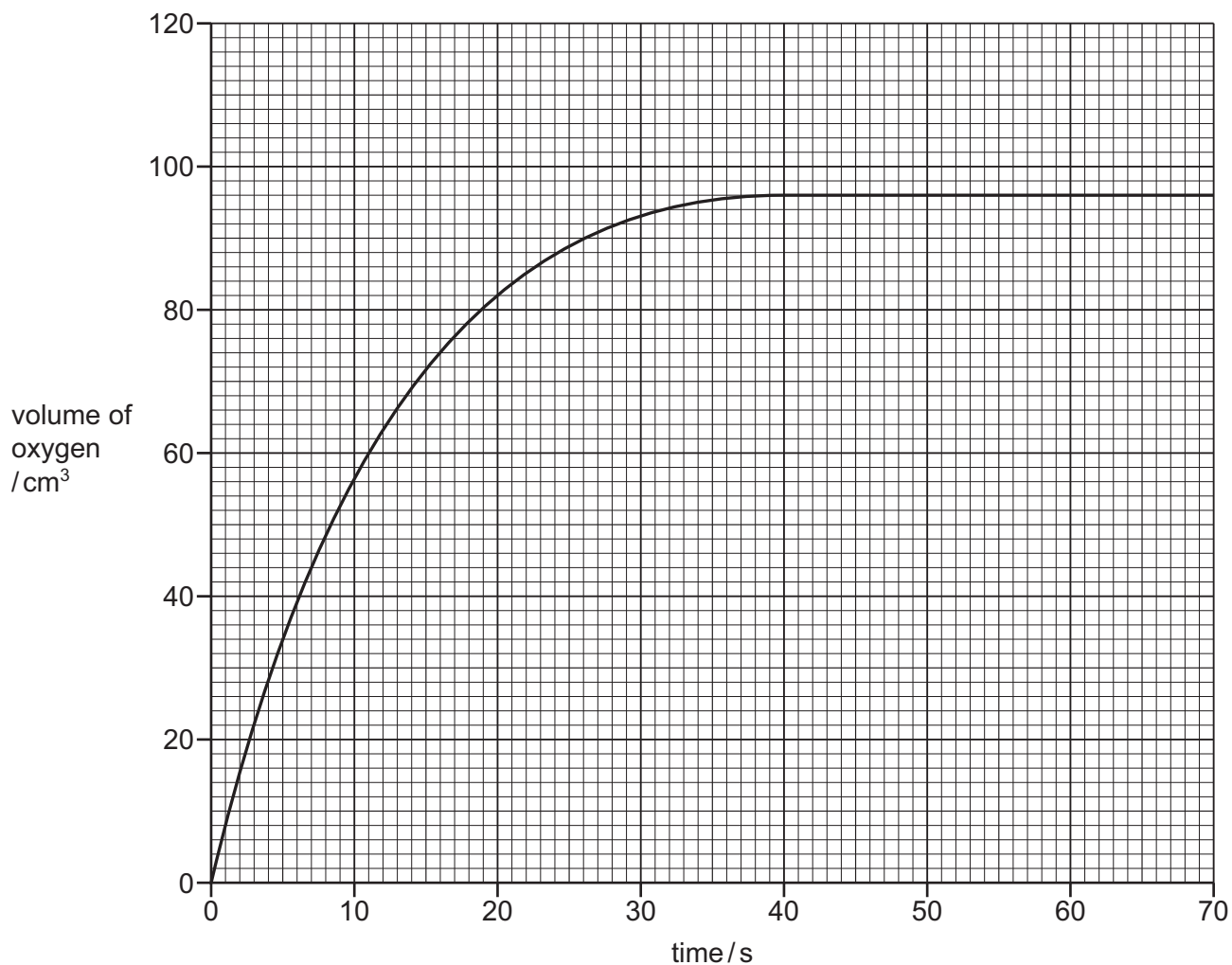
- (a) Complete the diagram to show a suitable method for collecting and measuring the volume of the oxygen.

Label your diagram.



[3]

(b) The graph shows how the volume of oxygen changes as the reaction proceeds.



Answer these questions using information from the graph.

(i) Describe how the rate of this reaction changes with time.

..... [1]

(ii) Deduce the time taken to collect 60 cm³ of oxygen.

time = s [1]

(iii) The experiment is repeated without using a catalyst.

Draw a line **on the grid** to show how the volume of oxygen changes with time when no catalyst is used.

All other conditions stay the same. [1]

(iv) Describe what effect an increase in temperature has on the rate of this reaction.

All other conditions stay the same.

..... [1]

(c) Identify which **one** of these elements is likely to act as a catalyst in chemical reactions.

Draw a circle around the correct answer.

C Mg Na Ni S [1]

(d) Describe a test for oxygen.

test

result

[2]

[Total: 10]

8 This question is about metals and compounds of metals.

- (a) Iron reacts with dilute hydrochloric acid to form an iron(II) salt and a gas which pops with a lighted splint.

Complete the word equation for this reaction.



[2]

- (b) Identify two correct statements about iron.

Tick **two** boxes.

Iron forms an alloy called steel.

☐

The commonest ore of iron is called bauxite.

☐

Iron is usually extracted from its ore by electrolysis.

☐

Iron is oxidised by carbon in the blast furnace.

☐

Both oxygen and water are needed for iron to rust.

☐

[2]

- (c) The table compares the reactions of four metals with warm water and with steam.

metal	reaction with warm water	reaction with steam
chromium	no reaction	slow reaction
copper	no reaction	no reaction
iron	very slow reaction	slow reaction
magnesium	very slow reaction	rapid reaction

Put the four metals in order of their reactivity.
Put the least reactive metal first.

least reactive	<div style="display: inline-block; width: 400px; border-bottom: 1px solid black; position: relative; height: 10px;"> <div style="position: absolute; right: -5px; top: -5px; width: 0; height: 0; border-left: 5px solid transparent; border-right: 5px solid transparent; border-bottom: 10px solid black;"></div> </div>	most reactive	

[2]

- (d) Crystals of cobalt(II) chloride, $\text{CoCl}_2 \cdot 6\text{H}_2\text{O}$, can be prepared by reacting excess cobalt(II) carbonate powder with dilute hydrochloric acid.

Describe how to prepare a sample of pure dry cobalt(II) chloride crystals after the reaction is complete.

In your answer describe how to:

- remove the excess cobalt(II) carbonate from the reaction mixture
- crystallise the cobalt(II) chloride
- dry the crystals.

.....

.....

.....

.....

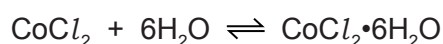
.....

.....

.....

..... [4]

- (e) A few drops of water were added to a sample of solid anhydrous cobalt(II) chloride, CoCl_2 . The equation for the reaction is shown.



- (i) State the meaning of the symbol $\rightleftharpoons$.

..... [1]

- (ii) State the colour change observed when water is added to anhydrous cobalt(II) chloride.

from to [2]

[Total: 13]

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The Periodic Table of Elements

Group																		
I	II											III	IV	V	VI	VII	VIII	
		<div>Key<div>atomic number atomic symbol name relative atomic mass</div></div>										<div>1Hhydrogen1</div>						
3Li lithium 7	4Be beryllium 9											5B boron 11	6C carbon 12	7N nitrogen 14	8O oxygen 16	9F fluorine 19	10Ne neon 20	
11Na sodium 23	12Mg magnesium 24											13Al aluminium 27	14Si silicon 28	15P phosphorus 31	16S sulfur 32	17Cl chlorine 35.5	18Ar argon 40	
19K potassium 39	20Ca calcium 40	21Sc scandium 45	22Ti titanium 48	23V vanadium 51	24Cr chromium 52	25Mn manganese 55	26Fe iron 56	27Co cobalt 59	28Ni nickel 59	29Cu copper 64	30Zn zinc 65	31Ga gallium 70	32Ge germanium 73	33As arsenic 75	34Se selenium 79	35Br bromine 80	36Kr krypton 84	
37Rb rubidium 85	38Sr strontium 88	39Y yttrium 89	40Zr zirconium 91	41Nb niobium 93	42Mo molybdenum 96	43Tc technetium —	44Ru ruthenium 101	45Rh rhodium 103	46Pd palladium 106	47Ag silver 108	48Cd cadmium 112	49In indium 115	50Sn tin 119	51Sb antimony 122	52Te tellurium 128	53I iodine 127	54Xe xenon 131	
55Cs caesium 133	56Ba barium 137	57–71 lanthanoids		72Hf hafnium 178	73Ta tantalum 181	74W tungsten 184	75Re rhenium 186	76Os osmium 190	77Ir iridium 192	78Pt platinum 195	79Au gold 197	80Hg mercury 201	81Tl thallium 204	82Pb lead 207	83Bi bismuth 209	84Po polonium —	85At astatine —	86Rn radon —
87Fr francium —	88Ra radium —	89–103 actinoids		104Rf rutherfordium —	105Db dubnium —	106Sg seaborgium —	107Bh bohrium —	108Hs hassium —	110Ds darmstadtium —	111Rg roentgenium —	112Cn copernicium —					116Lv livermorium —		

lanthanoids	57	La	lanthanum	139	58	Ce	cerium	140	59	Pr	praseodymium	141	60	Nd	neodymium	144	61	Pm	promethium	—	62	Sm	samarium	150	63	Eu	europium	152	64	Gd	gadolinium	157	65	Tb	terbium	159	66	Dy	dysprosium	163	67	Ho	holmium	165	68	Er	erbium	167	69	Tm	thulium	169	70	Yb	ytterbium	173	71	Lu	lutetium	175	
	actinoids	89	Ac	actinium	—	90	Th	thorium	232	91	Pa	protactinium	231	92	U	uranium	238	93	Np	neptunium	—	94	Pu	plutonium	—	95	Am	americium	—	96	Cm	curium	—	97	Bk	berkelium	—	98	Cf	californium	—	99	Es	einsteinium	—	100	Fm	fermium	—	101	Md	mendelevium	—	102	No	nobelium	—	103	Lr	lawrencium	—

The volume of one mole of any gas is 24 dm^3 at room temperature and pressure (r.t.p.).